

Semester Exam Formula Chart

$$\text{Percent Yield} \rightarrow \text{Percent Yield} = \frac{\text{Actual Yield}}{\text{Theoretical Yield}} \times 100\%$$

$$\text{Pressure Units} \rightarrow 1 \text{ atm} = 760 \text{ torr} = 760 \text{ mmHg} = 101.325 \text{ kPa}$$

$$\text{Temperature Units} \rightarrow K = ^\circ C + 273$$

$$\text{Boyle's Law} \rightarrow P_1 V_1 = P_2 V_2$$

$$\text{Charles' Law} \rightarrow \frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$\text{Combined Gas Law} \rightarrow \frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$\text{Ideal Gas Law} \rightarrow PV = nRT \quad \text{where } R = 0.821 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}}$$

$$\text{Dalton's Law of Partial Pressure} \rightarrow P_{\text{Total}} = P_1 + P_2 + P_3 \dots$$

$$\text{Concentration or Molarity} \rightarrow M = \frac{\text{mol}}{\text{L}}$$

$$\text{pH} \rightarrow \text{pH} = -\log[H^+] \text{ or } \text{pH} = -\log[H_3O^+]$$

$$\text{pOH} \rightarrow \text{pOH} = -\log[OH^-]$$

$$\text{pH vs pOH} \rightarrow \text{pH} + \text{pOH} = 14$$

$$\text{Concentration of an acid} \rightarrow [H^+ \text{ or } H_3O^+] = \text{antilog}^{-\text{pH}}$$

$$\text{Concentration of a base} \rightarrow [OH^-] = \text{antilog}^{-\text{pOH}}$$

$$\text{Heat} \rightarrow Q = m \cdot c \cdot \Delta T \quad c = 4.186 \text{ J/g}\cdot^\circ\text{C}$$

$$\text{Change in Enthalpy AKA Heat of Reactants} \rightarrow \Delta H = H_{\text{products}} - H_{\text{reactants}}$$

